

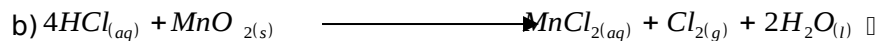
CHLORINE AND ITS COMPOUNDS

MARKING SCHEME

1. i) Anhydrous calcium chloride
 ii) A white ppt is formed
 HCl gas forms Cl^- ions solution which react with silver ions to form silver chloride which is insoluble
 OR
 $\text{HCl}_{(\text{aq})} + \text{AgNO}_{3(\text{aq})} \longrightarrow \text{HNO}_{3(\text{aq})} + \text{AgCl}_{(\text{s})} \checkmark 1\text{mk}$
 White ppt $\checkmark \frac{1}{2}$
 $\text{Cl}^-_{(\text{aq})} + \text{Ag}^+_{(\text{aq})} \longrightarrow \text{AgCl}_{(\text{s})}$

- 2.a) Potassium manganate VII ($\frac{1}{2}$)
 Manganese IV Oxide ($\frac{1}{2}$) **Any two 1mk**
 Lead IV Oxide ($\frac{1}{2}$)
 b) It is a bleaching agent

3. a) Oxidizing agent \square



4. a) $2\text{OCl}_{(\text{aq})} \longrightarrow 2\text{Cl}_{(\text{aq})} + \text{O}_{2(\text{g})} \checkmark 1$

- b) Time candle brightness $\checkmark 1$
 Colourless liquid (droplets) forms on walls $\checkmark 1$
 c) polypropene $\checkmark \frac{1}{2}$
 Making ropes $\checkmark \frac{1}{2}$
 Polytetrafluoroethene Tetrafluoroethene $\checkmark \frac{1}{2}$