

9.0.0 Salts M/S

1.
 - a) Conc. H_2SO_4 / H_2SO_4
 - b) Heat the solution to concentrate it.
Allow for crystals to form $\frac{1}{2}$ Filter $\frac{1}{2}$
 - c) Anhydrous Copper(II) sulphate/ $CuSO_4(s)$
 2.
 - a) To MgO , add excess HNO_3 , $\frac{1}{2}$ HCl or H_2SO_4 . Add $NaOH$ or KOH or NH_4OH to the mixture, $\frac{1}{2}$ Filter $\frac{1}{2}$ and dry $\frac{1}{2}$ the residue.
 - b) – Anti-acid (Treatment of acid indigestion)
- Making tooth past $\frac{1}{2}$
 3. Add excess lead (II) Oxide to dilute nitric (v) acid and filter to get lead (II) nitrate solution. Add sodium carbonate solution to lead (II) nitrate to precipitate lead (II) carbonate and wash with distilled water.
 4.
 - a) They became a white powder
 - b) Efflorescency
 5. Add water to sodium oxide to form sodium hydroxide solution. Bubble excess carbon (IV) oxide in sodium hydroxide solution to form sodium hydrogen carbonate. Heat sodium hydrogen carbonate solution to evaporate water.
 6. NH_4Cl decomposes on heating to produce NH_3 and $HCl(g)$. $NH_{3(g)}$ is lighter than $HCl_{(g)}$ hence diffuses faster and turns red-litmus to blue HCl is denser hence diffuses at a slower rate: changes blue litmus to red
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