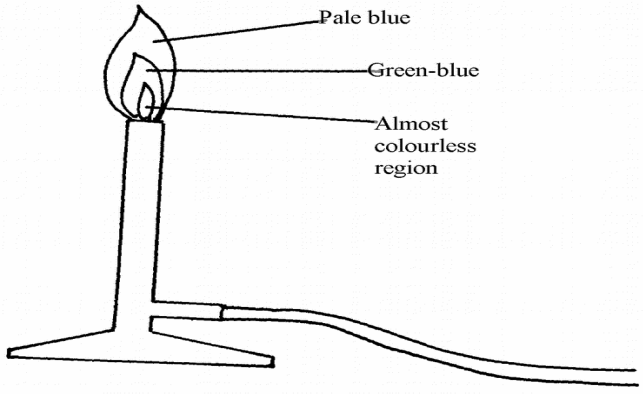
ANSWERS – CHEMISTRY
FORM ONE

Introduction to chemistry

1. a) F is place in the middle of the flame while G is placed at the upper region of the flame
   b) Non- luminous flame

2 a) The laboratory gas burns in excess oxygen
   OR burns completely or produces CO_2 and H_2O only
   - No unburnt carbon remains
   OR No soot is formed// Produced.

b)



3. a) a substance which when taken alters the body chemistry

   b) - alcohol
   - Tobacco

4. (a) A- Downward delivery /upward displacement of air
   B – Over water ✓ ½
   (b) A – Denser than air

5. (i) P – Haxane
   (ii) W – Water

6. Name – Mortar. √½
   Use – Holding solid substances being crushed. √½
   Name – Crucible √½
   Use – Holding solid elements being heated strongly. √½

7. T – has a very small hole which releases the gas in small quantities /in form of a jet.
   U – It is heavy for stability ✓1

8. (a) It is very hot. (1 mk) √1
   (b) The upper√1 part. Because all the gases undergo complete √1 combustion. √1 (2 mk)

9. The crystal dissolved√ ½ . Blue colour spreads in water √ ½ . The crystal broke up in smaller particles of copper (ii) sulphate and diffused in all direction

SIMPLE CLASSIFICATION OF SUBSTANCES

1.

(a) chromatography (1) [1]

(b) line drawn on diagram below origin (1) [1]

(c) does not interfere with results/owtte (1) [1]

2.

(a) A thermometer (1)

B beaker (1)

C tripod (1) [3]

(b) to cool / condense the vapour (1) [1]

(c) measure the boiling point (1) [1]

[Total: 5]

3.

(a) evaporation then condensation ✓ *if more than one box is ticked, award* 1

no mark

(b) water *accept 'distillate'* 1

(c) 100 *accept a temperature from 99 to 101* 1

(d) (i) a temperature above 15 but below 100 1

any **one** from

· it is heated by the water vapour

*accept 'vapour **or** steam **or** hot water' for water vapour*

· thermal energy **or** heat is transferred from the water vapour

accept 'it heats up'

***both** the answer and the correct explanation are required for the mark*

(ii) any **one** from 1

· it condenses *accept 'it makes condensation'*

· it changes to a liquid *accept 'it is liquid'*

accept 'it changes state'

accept 'it turns to water'

· its temperature falls *accept 'it cools'* 1

'it changes temperature' is insufficient

(e) any **one** from 1

· cold water replaces warm water

*accept 'it cools the water vapour better **or** more efficiently **or** quickly'*

accept 'it will work faster'

· the water stays cold **or** cooler

· there is a bigger difference in temperature between the water vapour and the water

· some vapour escapes in B

*accept 'it is cooled over a bigger length **or** for longer'*

accept 'more water is collected'

[7]

4.

(a) negligible/(approximately) $\frac{1}{2000}$; [Reject zero]

0;

1;

+1; 4

(b) (i) chlorine;

17 electrons; 2

(ii) 17; 1

(iii) A description to include:

1. (pale) green/yellow-green;

2. gas; 2

[9]

5.

i) I. SoapP1, sodiumP1 stearateP1

II. HydrolysisP1

ii) To precipitate P1soap/ solid L

6.

(a) (i) C 1

(ii) D 1

(iii) A and B *answers may be in either order* 1

***both** answers are required for the mark*

(iv) A and D *answers may be in either order* 1

***both** answers are required for the mark*

(v) C 1

(b) (i) the same *accept 'seven'* 1

(ii) a random, mixed arrangement of both types of molecule should be drawn with the molecules not touching each other 1

ACIDS, BASES AND INDICATORS

1.

(a)

liquid added	colour	acidic, alkaline or neutral?
water	purple	neutral
lemon juice	red	acidic
washing up liquid	blue	alkaline

3

(b) indicators *if more than one box is ticked, award no mark* 1

[4]

2.

(a) (i) any **one** from 1

· lemonade

· fruit juice *accept 'fruit **or** juice'*

(ii) any **one** from 1

· milk

· water

(b) (i) any **one** from 1

· blue

· dark blue

· purple

(ii) any **one** from 1

· it bubbled *accept 'it fizzed'*

· bubbles were formed *accept 'bubbles'*
accept 'effervescence'

(c) neutralisation ✓ 1
if more than one box is ticked, award no mark

[5]

3.

(a) (i) 7 *do **not** accept 'neutral'* 1

(ii) it was neutral ✓ *if more than one box is ticked,* 1
award no mark
consequential marking applies
accept 'it was acidic' if the answer to part (i) was less than 7
accept 'it was alkaline' if the answer to part (i) was greater than 7 and up to 14

(b) any **one** from 1

· it decreased **or** went down *accept 'it dropped to 5'*

· it became acidic **or** more acidic

(c) an alkali ✓ *if more than one box is ticked, award no mark* 1

[4]

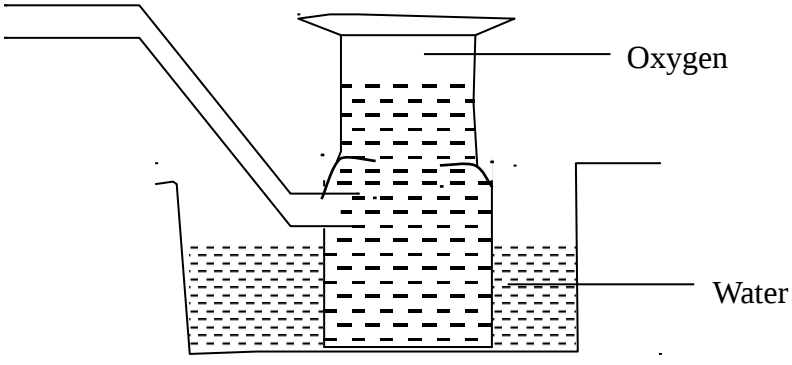
4.

(a) solution X = acid (1) Y and Z (both needed) are alkaline/alkali (1) [2]
(b) (i) barium chloride (nitrate) (solution) [1]
(ii) white (precipitate) (independent mark) [1]
accept milky/chalky
(iii) sulphuric acid [1]
accept correct formula where given but not hydrogen suphate
(c) (i) not enough of solution X had been added to react with all of solution Y (OWTTE)
(an understanding that sufficient acid must be added) [1]
(ii) the colour changed from pink to colourless [1]
(iii) neutralisation [1]

[Total: 8]

Air and combustion

1.



2. a) $3\ M_g + N_{2\,g} \_\_\_\_\_ M_{g\,3}\ N_{2\,g}$
b) Argon
- It is inert

3. a) Rust is hydrated iron (III) Oxide
b) - Electroplating
- Painting
- Oiling
- Galvanization
c) - Salts
- Acids

4. a) Moles of copper $^{8}/_{64}$ = 0.125 moles of Mg $^{3}/_{24}$ = 0.125Mg reacts with both O2 and N2 gases in the air while copper reacts with)2 only
There is greater change in the reaction with copper and smaller change in reaction with Mg

b) $CUO_{(g)} + H_2SO_{4(q)} \_\_\_\_\_ CUSO_{4(aq)} + H_2O_{(l)}$
Balanced

Chemical symbols correct
State symbols correct

5. a) Dust particles
b) They readily solidify hence may block the pipes
c) Argon

6. - Water rose up the test-tube to occupy the space of active air $\surd$½ which has been used in resting. $\surd$½
(a) Iron wool turned reddish – brown $\surd$½ due formation of red-oxide of iron $\surd$½ which is rust.

7. a)i)rusting occurred√ ½
ii) No rusting√ ½
b) In (i) iron is more reactive than copper hence undergoes corrosion√1
in (ii) zinc is more reactive than iron hence undergoes corrosion in place of iron√1

8. a) To remove any magnesium oxide coating from the surface of magnesium// To remove any oxide film on it

b) White solid which is magnesium oxide

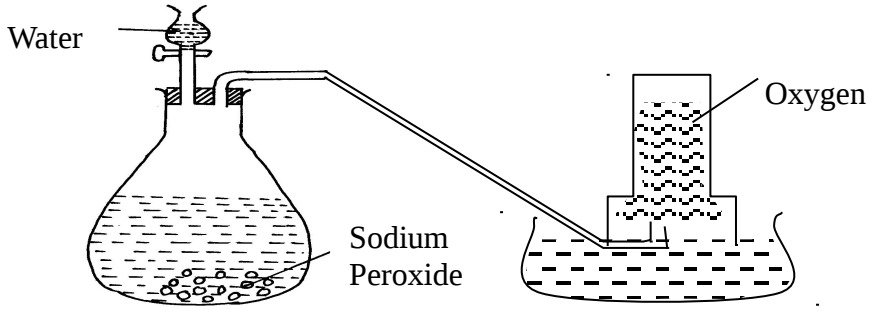
c) Increase in mass was due to oxygen which combined with magnesium

d) $2Mg(s) + O_{2(g)}$ _____ $2MgO(s)$
Penalize ½ for wrong or missing state symbols

e) The filtrate is magnesium hydroxide which is an alkaline
Red litmus paper changed blue, but blue litmus paper remained blue

9. (a) So that they may stick to the gas Jar to prevent them from falling into water when the gas jar is inverted
(b) Iron filings turned to reddish brown because they reacted with oxygen in presence of moisture to form rust.
- The level of water inside the gas jar rise so as to occupy the volume initially occupied by part of air used up for rusting
(c) - Air is made up of two parts; - the active part that is necessary for rusting and the inactive part that is not used for rusting
- oxygen is the active part of air

(d)



- Neat diagram-
- correct method of collection

(e) - For cutting and welding metals
- Rocket fuel
- Mountain climbing
- Sea diving
- Used in explosions (any two)

10. a) To remove any magnesium oxide coating from the surface of magnesium// To remove any oxide film on it

b) White solid which is magnesium oxide

c) Increase in mass was due to oxygen which combined with magnesium

d) $2Mg(s) + O2(g) \_\_\_\_\_ 2MgO_{(s)}$
Penalize ½ for wrong or missing state symbols

e) The filtrate is magnesium hydroxide which is an alkaline
Red litmus paper changed blue, but blue litmus paper remained blue

11. (i) Oxygen
(ii) Sodium hydroxide is a strong base
(iii) Slightly soluble in water

12. (i) White fumes form in the gas jar which disappear after sometime.
- The level of water rises in the gas jar.

(ii) $P_{(s)} + O_{2(g)} \longrightarrow P_2O5_{(s)}$
$P_2O_{(s)} + 3H_2O_{(l)} \longrightarrow 2H_4PO_{4(aq)}$

(iii) Magnesium react with oxygen and nitrogen hence greater of fraction of air is used.

(iv) (a) Blue litmus changed to red as remained red. The solution was acid due to phosphoric
(b) Red litmus changed to blue as blue remained blue due to formation of basic magnesium hydroxide ammonia solution.

(v) – Pass air over conc. KOH / NaOH to absorb CO_2
- Pass the remaining gases over hot copper solid which reacts with oxygen.
- Collect the remaining gas over water. The gas is mainly nitrogen.

13. a) i) $3Mg(s) + N_2(g) \longrightarrow Mg_3N_2(s)$ √1
ii) Gas with√1 choking irritating smell.
Mg_3N_2 reacts with water to form ammonia √1 gas.
iii) It remains blue. √½ Ammonia gas is alkaline. √½

14. (a) (i) Phosphorous
(ii) - Do not react with water when being inserted into the tube
- reacts with oxygen when exposed to air.

(b) $4P(s) + 3O_{2(g)} \longrightarrow 2P_2O_{3(s)}$
or $4P(s) + SO_{2(g)} \longrightarrow 2P_2O_{5(s)}$

(c) (i) $\frac{Y - X}{} \times 100$

y

(ii) – Wrong reading of volume

- Phosphorous can go off before complete combustion

(d) (i) – Red litmus paper no effect

- Blue litmus paper turns red due to formation of phosphoric acid/phosphorous (V) Oxide whish is an acidic oxide

(ii) – Oxygen

(iii) – Burning of candle

- Use of pyrogallol

- Rusting of iron fillings

15. i) $P_{4(g)} + 5O_{2(g)} \_\_\_\_\_ 2P_2O_{5(s)}$

// $P_{4(s)} + 3O_{2(g)} \_\_\_\_\_ 2P_2O_{3(g)}$ Anyone √ 1 mark

ii) Phosphorous (v) or (iii) oxide formed is an acidic Oxide which dissolves in water to form a strong acidic solution of phosphoric acid whose PH is 2

16. (a) – Iron nails turns brown.

- Water rises up the delivery tube/water level drops in the trough (any ½mk) ✓ ½

Explanation: Oxygen has been used up in rusting of iron nails hence water rises up to take the place of oxygen ✓1

(b) $4Fe_{(s)} + 3O_{2(g)} + 2H_2O_{(l)} \longrightarrow 2Fe_2O_3.2H_2O_{(s)}$ ✓ 1

(accept a balanced chemical equation)

17. a) $FeCO_{3\,(s)} \longrightarrow Fe\,O_{(s)} + CO_{2(g)}$

$Fe(s) + 4\,H_2O_{(g)} \longrightarrow Fe\,O_{4\,(s)} + 4H_{2\,(g)}$

Or

$2\,Fe_{(s)} + 2O_{2(g)} \longrightarrow Fe_3\,O_{4(s)}$

b) $Fe_3O_{4(s)} + 8H^{+}_{(aq)} \longrightarrow 4H_2O_{(l)} + 2\,Fe^{3+}_{(aq)} + Fe^{2+}_{(aq)}$

18. a) N_2O 1 (Nitrogen (I) oxide) – Denitrogen Oxide.

b) K_2O 1 (Potassium oxide)

c) Al_2O_3 (Aluminium oxide)

19. a) water √1

b) $2Na_2O_{2(S)} + 2H_2O_{(L)} \longrightarrow 4NaOH_{(aq)} + O2_{(g)}$ √1 mk

Penalize ½ - wrong missing state symbols

Water and hydrogen

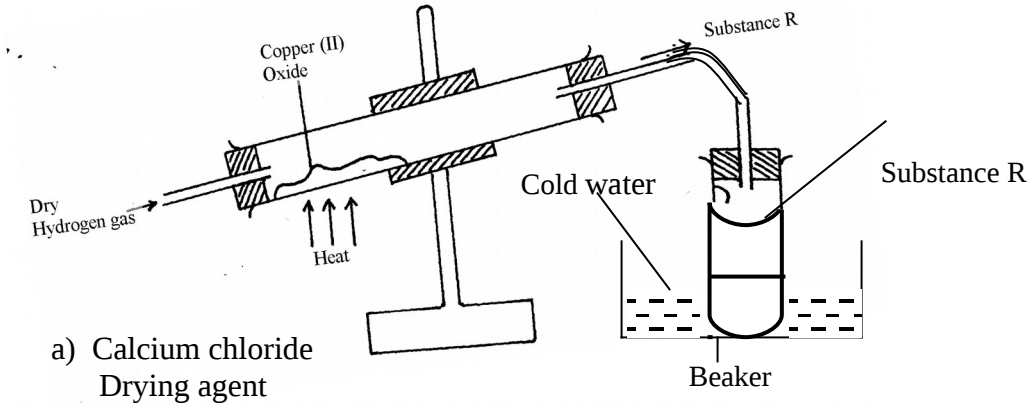
1. (a) Aluminium is above hydrogen in the reactivity series of elements

(b) (i) The reaction is too exothermic that alot of heat is produced causing ignition of hydrogen in presence of oxygen

(ii) $K_{(s)} + H_2O_{(g)} \longrightarrow KOH_{(aq)} + H_{2(g)}$

$H_{2(g)} + O_{2(g)} \longrightarrow H_2O_{(g)}$

2.



3. a) Calcium chloride
Drying agent

b) $2H_{2(g)} + O_{2(g)}$ _____ $2H_2O_{(g)}$

4. (i) Steam

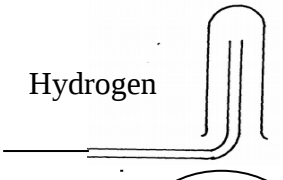
(ii) $Mg_{(s)} + H_2O_{(g)}$ $MgO_{(s)} + H_{2(g)}$ ✓ 1

(iii) Gas P is passed through the combustion tube before heating is commenced

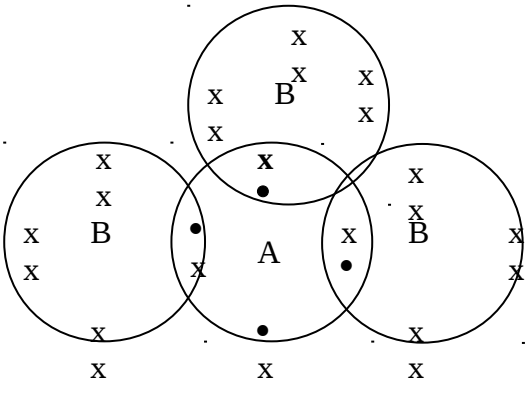
5. a) $2H_{2(g)} + O_{2(g)} \longrightarrow 2H_2O_{(l)}$ √1

b) – Turns anhydrous white paper √½ copper (II) sulphate into blue. √½ Or
- Turns anhydrous blue √½ cobalt (II) chloride into pink. √½

6. a)



b)reverse steam√1



7. (a) N

(b) $4H_2O_{(g)} + 3Fe_{(s)} \longrightarrow Fe3O4_{(s)} + 4H_{2(g)}$ (Not balanced 0mk)