CHEMISTRY PAPER 1 MARKING SCHEME
FORM 3 END-TERM I EXAM
2017

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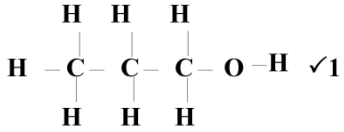
1.a)Ionization energy is the energy required to remove an election from ✓ 1 atom in gaseous state while electron affinity is the energy required by an atom ✓ 1 to require an electron in gaseous state.

b)B is higher / greater ✓ ½ than A because A is smaller atom therefore its nuclear attract electrons strongly ✓ ½

2. i)



ii) Dehydration ✓1

iii) C_6H_{12} ✓1

3.a) Covalent bond ✓1

b) Giant atomic structure ✓1

c)Hard ✓1 high density // high melting points Silicon and oxygen ✓1 atoms are compactly held by strong covalent bonds throughout its structure

4. Add ✓ ½ excess lead (II) carbonate to dilute nitric (V) acid

- Filter ✓ ½ to remove excess ✓ ½ unreacted lead (II) carbonate
- Add ✓ ½ dilute hydrochloric acid to the filtrate
- Filtre ✓ ½ and dry✓ ½ the residue

5a) The rate of diffusion of a gas is inversely ✓1 proportional to the square root of its density provide temperature and pressure are kept constant

$$\frac{RNH_3}{RxH_3}\sqrt{\frac{N_xH_3}{M_NH_3}} = 1.41$$

$$\sqrt{\frac{M_xH_3}{17}} = 1.41^2$$

$$\frac{M_xH_3}{17} = 1.41^2$$

b) $M_xH_3 = 17 \times 1.14^2$

$= 33.7977$

$RAM\ of \times = 33.7977 - 3$

$= 28.7977$

6.

Fe	O
7 $7/56$ $\frac{0.125}{0.125}$	3 $3/6$ $\frac{0.1875}{0.125}$ ½ 1.5

2:3

Fe_2O_3

$2\ Fe_2O_{3(s)} + 3\ C_{(s)} \longrightarrow 4\ Fe_{(s)} + 3\ CO_{2(g)}$ ✓

7a)A: Sublimation✓

B: Deposition✓

8.a)Measure of acidity or basically of an aqueous solution✓ (1mk)

b) B✓ (1mk)

c) 10 (1mk)

10. (a) (i) Ionic bond ✓1

(ii) Covalent bond ✓1

(b) T ✓ ½ and W ✓ ½ (a)

11. $2H_2O_{2(aq)} \longrightarrow 2H_2O_{(l)} + O_{2(g)}$ ✓1

- U.B eqn. – zero mk
- Penalise ½ mk for wrong or missing s.s

(b) Manganese (IV) oxide ✓1

(c) - Used in welding and cutting metals as oxyacetyline/ oxyhydrogen.
- Used to remove Iron impurities during steel making (Any 1 x 1mk) a)

12.a)On the diagram (left hand electrode)

b) $Pb_{(s)} \longrightarrow Pb^{2+}_{(aq)} + 2e$ $2e^- \longrightarrow Pb^{2+}_{(l)} + 2e^- \longrightarrow Pb_{(s)}$

c) Extraction of metals

13.(a) CO_2

(b) $2NaHCO_{3(s)} \longrightarrow Na_2CO_{3(s)} + CO_{2(g)} + H_2O_{(g)}$

(c) making glass, softening hard water

14.(i) Yellow lead (II) oxide turned to red then grey.

(ii) I. $H_{2(g)} + PbO_{(s)} \longrightarrow H_2O_{(l)} + Pb_{(s)}$

II. $2H_{2(g)} + O_{2(g)} \longrightarrow 2H_2O_{(l)}$

(iii) Reducing properties of hydrogen

Combustion nature of hydrogen

15.(a) Physical change ½ mk

(b) Chemical change

(c) physical change

(d) chemical change

16.a) Rusting

b)has water of crytalization

c) painting,electroplating,anodizing,Gavaization

16.(i) U

(ii) Molecular

(iii) X is smaller than W

17.. i) a) D 1mark

b) C 1mark

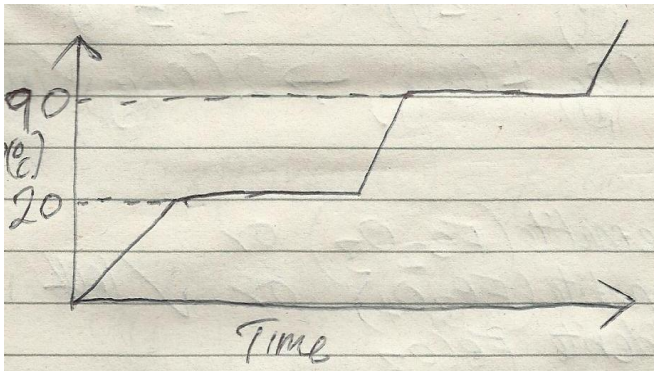
ii) B 1mark

19..a) A – fumes of colourless gas observed. Green solid turns black

B – White precipitate is observed 1mark

b) $CuCO_3$ (g) $\longrightarrow$ $CUO(s) + CO_2(g)$ ✓1

20.



21i) A: carbon reacts with excess air to form carbon(iv) oxide ½ mark

$C(s) + O_2$ (g) $\longrightarrow$ $CO_2(g)$ ½ mark

B: Carbon (iv) oxide is reduced to

Carbon (ii) oxide by hot carbon ½ mark

$C(s) + CO_2$ (g) $\longrightarrow 2CO(g)$ ½ mark

ii) Carbon (iv) oxide causes global warming 1mark

22.

	Na_2SO_4	H_2O	
Mass	1.42	1.8✓	½
RFM	142	18	
Moles	$\frac{1.42}{142}$	$\frac{1.8}{18}$ ✓	½
Divide by smallest No.	$\frac{0.01}{0.01}$	$\frac{0.1}{0.1}$ ✓	½
	1	10	

X = 10 ✓ ½

(2 mks)

23.(a) (i) S_{16} = 2.8.6 (1 mk)

(ii) S_{12} = 2.8.2 (1 mk)

(b) (i) Neutron – 14 (1 mk)

(ii) Electron - 10 (1 mk)

24.(i) At constant temperature the volume is inversely proportional to the pressure Formula

$V d \frac{1}{d}$ (1 mk)

(ii) $P_1V_1 = P_2V_2$ ✓ ½

$12 \text{ X}1 = 2.5 \text{ X } V_2$ ✓ ½

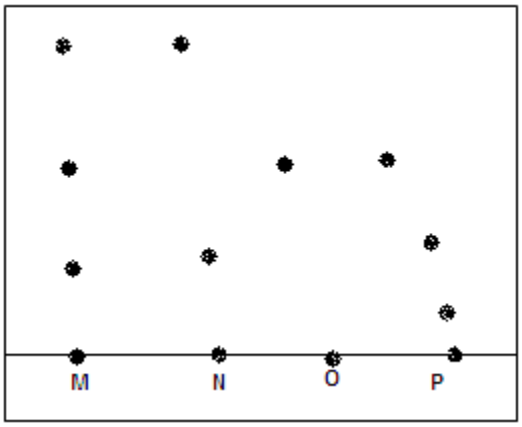
$V_2 = \frac{12\,X1}{2.5}$ ✓ ½ = 4.8 litres ✓ ½ (2 mks)

25a) Sample I is a pure substance since pure substance have a sharp melting and boiling points. (1mk)

Sample II is impure since the melting point is lower than that of a pure substance and its boiling point is higher than that of pure substance which is characteristic phenomena of an impure substance. (1mk)

b) Since ice causes skidding, common salt becomes an impurity to water (ice) causing it to melt at a lower temperature. .(1mk)

26a)



b) M has N and O (1)

a) P

27.a) X- fractionating column (1mk)

Y- Liebig condenser

b) to condense back the component of higher boiling point. (1mk)

c) shown on the diagram (1mk)