

NATIONAL EXAMINATIONS

FORM 4 MARCH TRIAL 2017 CHEMISTRY

CONFIDENTIAL

CONFIDENTIAL TO ALL SCHOOLS FOR CHEMISTRY TEACHERS

INSTRUCTIONS TO SCHOOLS

The information contained in this paper is to enable the head of the school and the teacher in charge of chemistry to make adequate preparations for this year's chemistry practical mock examination. **NO ONE ELSE** should have access to this paper or acquire knowledge of its contents. Great care must be taken to ensure that the information herein does not reach the candidates either directly or indirectly. Chemistry teachers **SHOULD NOT** perform any of the experiments in the same room as the candidates or make the results of the experiments available to the candidates or give any other information related to the experiments. Doing so will constitute an examination irregularity which is punishable.

EACH CANDIDATE WILL REQUIRE:

- ❖ One burette 0 – 50cm³
- ❖ One pipette 25cm³
- ❖ About 80 cm³ of solution 2M HCl
- ❖ About 100 cm³ of solution 0.1 NaOH
- ❖ About 80 cm³ of 0.05M H₂SO₄
- ❖ Two conical flasks (250cm³)
- ❖ One complete stand
- ❖ 100cm³ measuring cylinder
- ❖ 10 cm³ measuring cylinder
- ❖ Means of labeling
- ❖ Test-tube rack
- ❖ 5 clean dry test tubes
- ❖ White tile
- ❖ One 250cm³ beaker
- ❖ One spatula
- ❖ About one spatulaful solid E
- ❖ About one spatulaful solid M

ACCESS TO:

- ❖ Distilled water
- ❖ Phenolphthalein indicator supplied with a dropper
- ❖ 2M Lead (II) Nitrate supplied with a dropper
- ❖ 2M Hydrochloric acid supplied with a dropper
- ❖ 1M Barium chloride solution supplied with a dropper
- ❖ 2M aqueous Ammonia supplied with a dropper

NOTE

1. Solid M is lead (II) sulphate (PbSO₄)
2. Solid E is zinc chloride (ZnCl₂)