NAME: …… INDEX NO: ……

SCHOOL: …… DATE : ……

CANDIDATE'S SIGNATURE:……..

233/3
CHEMISTRY
PAPER 3
PRACTICAL
JULY / AUGUST 2014
TIME: 2 ¼ HOURS

NANDI CENTRAL SUB-COUNTY JOINT EVALUATION 2014

Kenya Certificate of Secondary Education (K.C.S.E.)
CHEMISTRY
PAPER 3
TIME: 2 ¼ HOURS

INSTRUCTIONS TO CANDIDATES

- Write your Name and Index Number in the spaces provided above.
- Answer all the questions in the spaces provided.
- You are not allowed to start with apparatus for the first 15 minutes of the 2¼ hours allowed for this paper. This time is to enable you to read the question paper and make sure you have all the apparatus and chemicals that you may need.
- ALL working must be clearly shown.
- Mathematical tables and electronic calculators may be used.

FOR EXAMINER'S USE ONLY

QUESTIONS	MAX SCORE	CANDIDATE'S SCORE
1	22	
2	12	

3	06	
TOTAL	**40**	

Procedure II

(a) (i) Transfer the contents of the boiling tube into a 250ml volumetric flask. Rinse both the boiling tube and the thermometer with distilled water and add the water to the volumetric flask. Add more distilled water to make up to the mark. Label this solution A. Fill a burette with solution A. Using a pipette and a pipette filler, place 25.0cm^3 of solution B into a conical flask. Add three (3) drops of phenolphthalein indicator. Titrate the solution A with solution B until the pink colour fades. Record your readings in Table 2. Repeat the titration two more times and complete Table 2.

Table 2

	I	**II**	**III**
Final burette reading			
Initial burette reading			
Volume of solution **A** used (cm^3)			

(ii) Calculate the:-

(a) Average volume of solution A used. (1mk)

(b) Number of moles of sodium hydroxide in 24cm3 solution pipette. (1mk)

(c) Number of moles of A in the average volume of solution given that solid a is a dibasic acid. (1mk)

(d) Relative formula mass of A. (3mks)

1 ONLY

1. You are provided with:-
   - Solid T, hydrated ethanedioc acid $H_2C_2O_4.nH_2O$.
   - Solution Q, a 0.2m solution of sodium hydroxide.

   You are required to determine:
   (i) Solubility of solid T.
   (ii) The value of n is the formula $H_2C_2O_4.nH_2O$.

   Procedure I
   (i) Fill the burette with distilled water.
   (ii) Place solid T in the boiling tube.
   (iii) Transfer $4cm^3$ of distilled water from the burette into the boiling tube containing solid T. Heat the mixture while stirring with the thermometer to a temperature at which crystals start to form in the table 1 below.
   (iv) Add a further $2cm^3$ of distilled water from the burette to the mixture. Repeat the procedure (iv) above and record the crystallization temperature. Complete the table I below by adding the volumes of distilled water as indicated.
   (Preserve the contents of the boiling tube for procedure II)

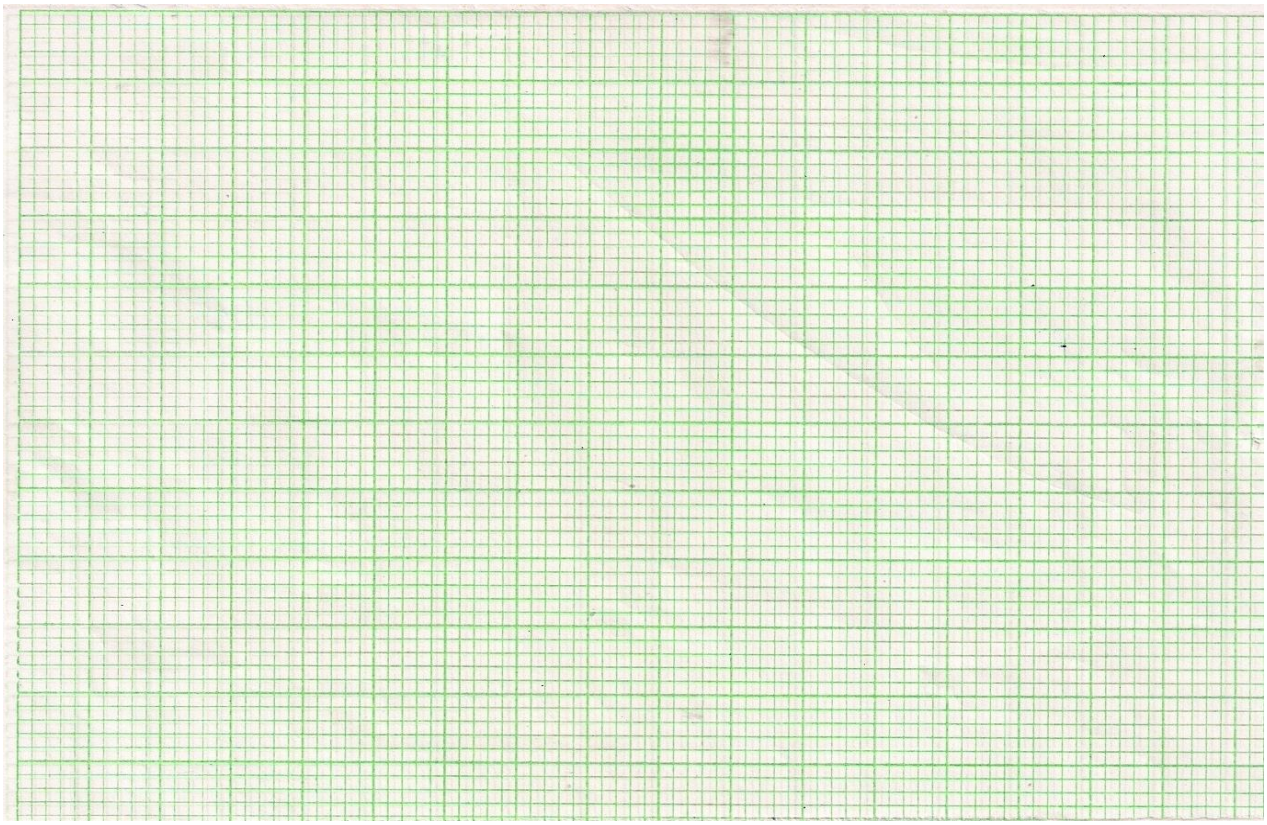
   TABLE I

Volume of distilled water in boiling tube	**Crystallization temperature**	**Stability of solid T in 100g / water**
4		

6		
8		
12		

(6mks)

(a) On the grid provided, plot a graph of solubility of solid T (y-axis) against crystallization temperature. (3mks)



From the graph determine:

(i) Solubility of T at 55^0C (1mk)

(ii) The temperature at which 80g of T dissolve in 100g of water. (2mks)

Procedure II

- Transfer the contents of the boiling tube in procedure I to a clean 200ml volumetric flask. Add distilled water to the mark. Label the resulting solution T.
- Fill the burette with solution T. Pipette $25cm^3$ of Q into a clean 200ml conical flask. Add 3 drops of phenolphthalein indicator.
- Titrate T against Q to an accurate end point. Record your results in the table II below.
- Repeat the experiment two more times and complete the table II below.

Table II

	I	II	III
Final burette reading cm^3			
Initial burette reading cm^3			
Volume of T used cm^3			

(4mks)

Calculate:

(a) Average volume of T used. (1mk)

(b) (i) Moles of Q used. (1mk)

(ii) Moles of T used. (1mk)

(iii) Concentration of T in molar per dm^3. (1mk)

(c) Determine the value of n in the formula $H_2C_2O_4.nH_2O$. (2mks)

2. You are provided with solution D. You are required to carry out the tests on solution and record your observations and inferences in the space provided.

(i) To about $2cm^3$ of solution D, add 3 drops of potassium iodide solution.

Observations	Inference

(1mk) (1mk)

(ii) To the remaining portion in the boiling tube add 5cm3 of dilute hydrocholic acid and warm. Leave it to cool and filter.

Observation	Inference
(1mk)	(1mk)

Divide the filtrate into two portions.

(iii) To one portion, add sodium hydroxide drop-wise until in excess.

Observation	Inference
(1mk)	(1mk)

(iv) To 2nd portion, add aqueous ammonic drop-wise till in excess.

Observation	Inferences
(1mk)	(1mk)

3. You are provided with solid R. Carry out the tests below and record your observations and inferences in the spaces provided.

(i) Place one third of solid R on a metallic spatula. Burn it in a non-luminous flame of the Bunsen Burner.

Observation	Inference
(1mk)	(1mk)

(ii) Place the remaining solid in a test-tube. Add about $6cm^3$ of distilled water and shake the mixture well.

Observation	Inference
(1mk)	(1mk)

Divide the solution into 2 portions.

(I) To about $2cm^3$ of the solution, add 1g of solid A; sodium hydrogen carbonate.

Observation	Inference
(1mk)	(1mk)

(II) To about $1cm^3$, add 3 drops of acidified chromate (vi) and warm.

Observation	Inferences
(1mk)	(1mk)

(III) In another $2cm^3$, add 2 drops of acidified potassium manganate (vii).

ANSWERS:

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